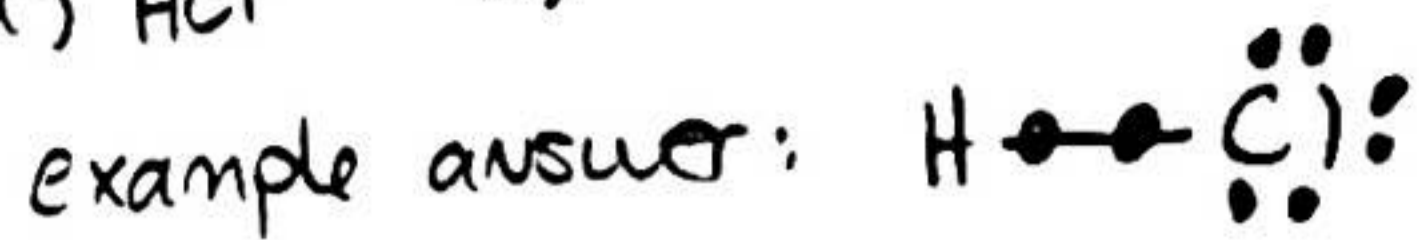


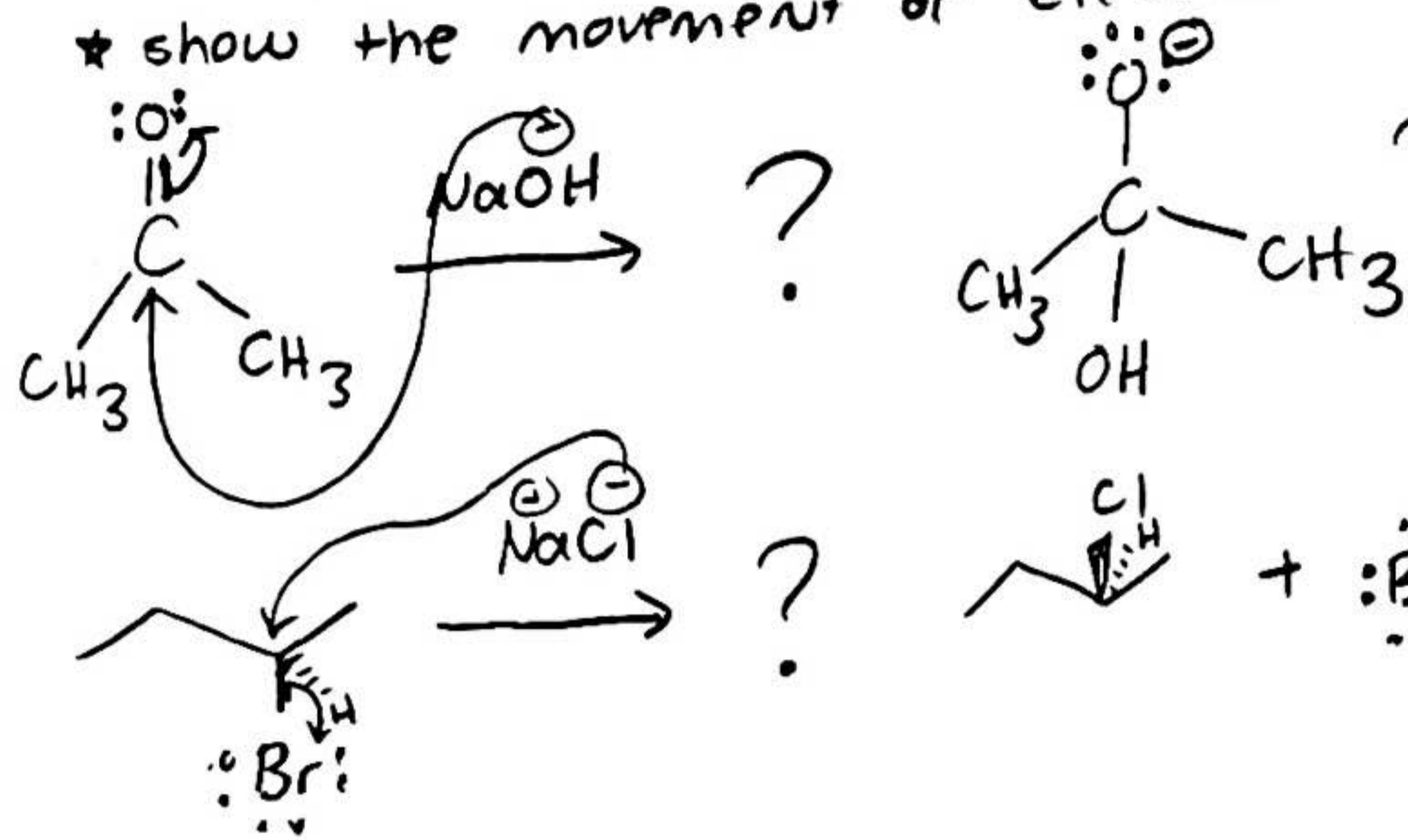
Draw the following Lewis structures.

SHOW ALL ELECTRONS!!!

- a) HCl b) HBr c) HPO_4^{2-} d) CH_3Br e) CO_2

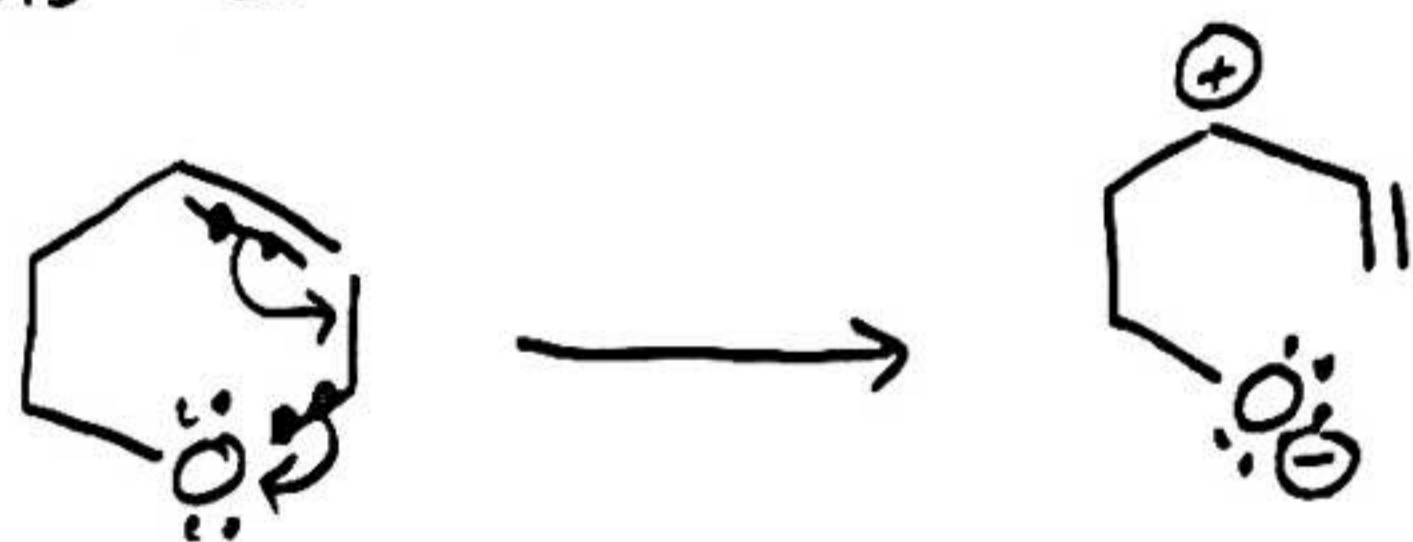


For the following reactions draw the products
 * show the movement of electrons (arrows)

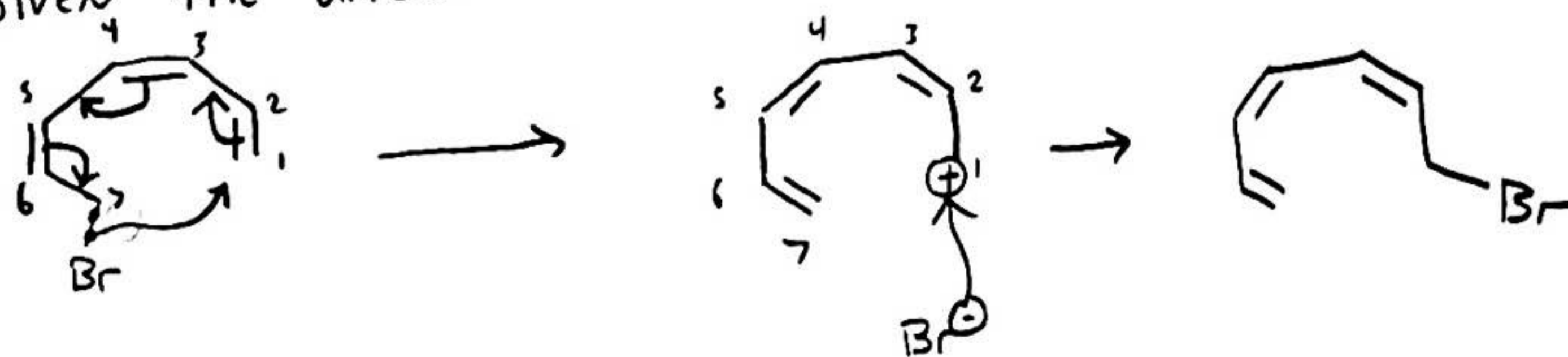


draw the products
 it is more likely that an acid base reaction will occur here in real life but for what you need to know right now the OH^- acts as a nucleophile
 (General concept for this point)
 * email me if you would like to know about the acid base rxn or actually does come to my session
 this BUT YOU DO NOT NEED TO KNOW THIS RIGHT NOW $\text{H}_2\text{O} + \text{H}^+ \rightarrow \text{H}_3\text{O}^+$

For the following draw arrows that properly produce the products shown:

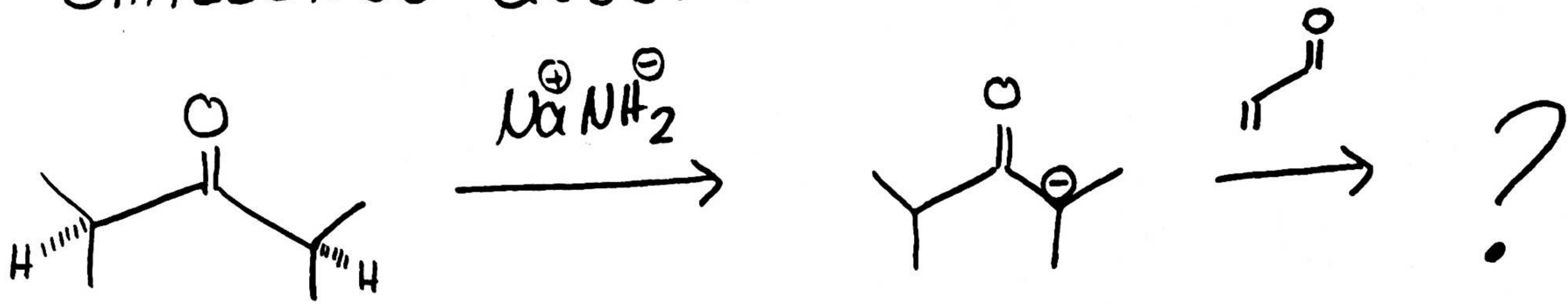


Given the arrows draw the product



PERFORMED

CHALLENGE QUESTION OF THE DAY



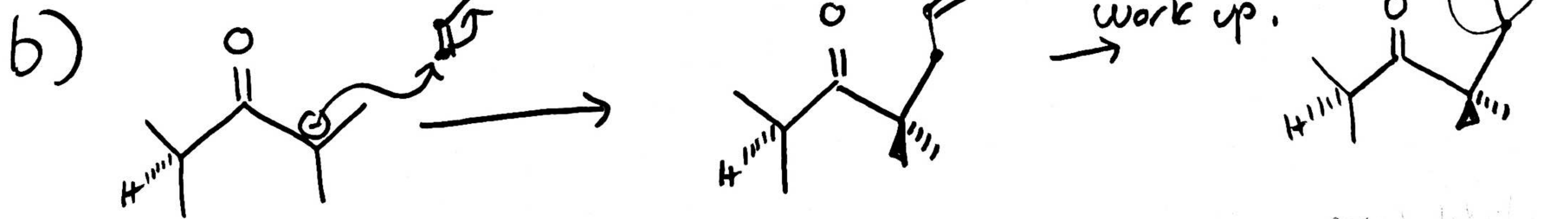
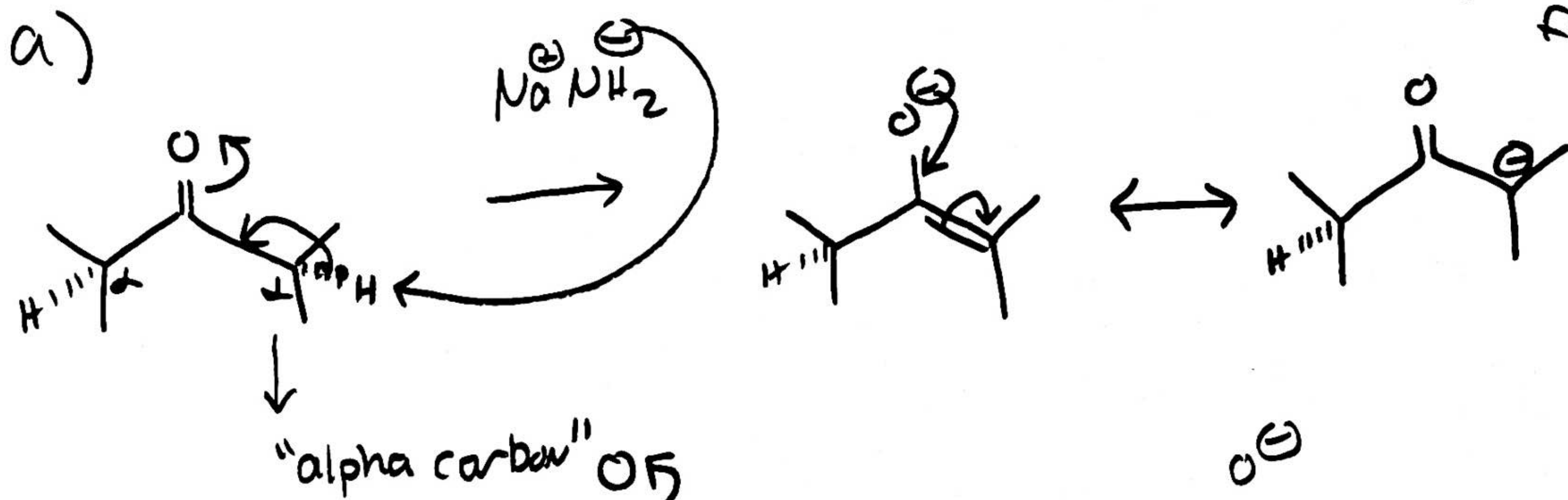
a) show how the negative charge got there.

show ALL intermediates

b) Draw the product. SHOW ALL ARROWS that formed the product.

(* may be helpful to review what a base is.)

a base is a hydrogen acceptor \therefore steals hydrogens from ACIDIC things. like alpha hydrogens on alpha carbons



most likely to be the product

THIS QUESTION IS WAY HARDER THEN ANYTHING YOU NEED TO KNOW. IF YOU UNDERSTAND THIS GREAT JOB!